

## Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write *true* or *false*.

- \_\_\_\_\_ 1. Binary molecular compounds are generally composed of a metal and a nonmetal.
- \_\_\_\_\_ 2. The second element in the formula of a binary compound is named using the suffix *-ite*.
- \_\_\_\_\_ 3. The prefix *tetra-* indicates three atoms.
- \_\_\_\_\_ 4. The prefix *hexa-* indicates six atoms.
- \_\_\_\_\_ 5. In naming the first element in a formula, the prefix *mono-* is not used.
- \_\_\_\_\_ 6. For binary acids, the hydrogen part of the compound is named using the prefix *hydro-*.
- \_\_\_\_\_ 7. An oxyacid contains only two elements.
- \_\_\_\_\_ 8. If the name of the anion of an oxyacid ends in *-ate*, the acid name contains the suffix *-ous*.

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

### Column A

- \_\_\_\_\_ 9. CO
- \_\_\_\_\_ 10. CO<sub>2</sub>
- \_\_\_\_\_ 11. H<sub>2</sub>CO<sub>3</sub>
- \_\_\_\_\_ 12. NH<sub>3</sub>
- \_\_\_\_\_ 13. N<sub>2</sub>O<sub>4</sub>
- \_\_\_\_\_ 14. HNO<sub>2</sub>
- \_\_\_\_\_ 15. HNO<sub>3</sub>
- \_\_\_\_\_ 16. HBr
- \_\_\_\_\_ 17. HBrO<sub>3</sub>

### Column B

- a. hydrobromic acid
- b. dinitrogen tetroxide
- c. carbon monoxide
- d. nitrous acid
- e. ammonia
- f. nitric acid
- g. carbonic acid
- h. bromic acid
- i. carbon dioxide