Section 13.3 Liquids and Solids

In your textbook, read about liquids and solids.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

_____ **1.** The constant *motion* of the particles in a liquid causes the liquid to take the shape of its container.

_____ **2.** At room temperature and one atmosphere of air pressure, the density of a liquid is much *greater* than that of its vapor.

_____ **3.** Liquids are not easily compressed because their particles are *loosely* packed.

_____ **4.** A liquid is less fluid than a gas because *intramolecular* attractions interfere with the ability of particles to flow past one another.

_____ **5.** Liquids that have stronger intermolecular forces have *higher* viscosities than do liquids with weaker intermolecular forces.

_____ **6.** The viscosity of a liquid *increases* with temperature because the increased average kinetic energy of the particles makes it easier for the particles to flow.

_____ **7.** Liquids that can form hydrogen bonds generally have a *high* surface tension.

_____ **8.** A liquid that rises in a narrow glass tube shows that the adhesive forces between the particles of the liquid and glass are *greater* than the cohesive forces between the particles of the liquid.

_____ **9.** Solids have a definite shape and volume because the motion of their particles is limited to *vibrations* around fixed locations.

_____ **10.** Most solids are *less* dense than liquids because the particles in a solid are more closely packed than those in a liquid.

_____ **11.** Rubber is *a crystalline* solid because its particles are not arranged in a regular, repeating pattern.